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| http://www.lyc-hoche-versailles.ac-versailles.fr/IMG/png/logo_ac-versailles_transparent_hd.png | **ÉVALUATION FORMATIVE** | Sciences Physiques  CAP |

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| **GRILLE NATIONALE D’ÉVALUATION EN MATHÉMATIQUES ET  EN SCIENCES PHYSIQUES ET CHIMIQUES** | | |
| NOM et Prénom : | Diplôme préparé : CAP | Séquence d’évaluation n° |

**Liste des capacités, connaissances et attitudes évaluées**

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| **Capacités** | * Sécurité : prévention des risques chimiques et électriques * Chimie 2 : acidité, basicité, pH |
| **Connaissances** | * Lire et exploiter les informations données sur l’étiquette d’un produit chimique de laboratoire ou d’usage domestique. * Mettre en œuvre les procédures et consignes de sécurité établies. * Reconnaître le caractère acide, basique ou neutre d’une solution. |

**Évaluation**

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| **Compétences** | **Capacités** | **Questions** | | **Appréciation du niveau d’acquisition** | | | |
| **A** | **ECA** | **NA** | |
| **S’approprier** | L’élève donne la signification du pictogramme.  L’élève identifie les dangers du produits par lecture de l’étiquette. | **Appel 1 et 2** | A |  |  | |  |
| B  C |  |  | |  |
|  |  |  | |  |
| **Analyser**  **Raisonner** | L’élève  -choisit une proposition, et propose une expérience  L’eau de javel est : ❑ acide neutre ❑ basique  Lorsqu’on dilue une solution d’eau de javel, le pH :  ❑ augmente ❑ reste inchangé ❑ diminue | I-Hypothèse |  |  | |  |
| I-Méthode de résolution |  |  | |  |
| II- Hypothèse |  |  | |  |
| II- Méthode de résolution |  |  | |  |
|  |  |  | |  |
| **Réaliser** | L’élève met en œuvre les précautions de sécurité  L’élève met en œuvre le protocole de mesure de pH en respectant le protocole  L’élève met dans l’ordre le protocole de dilution au dixième  L’élève réalise la manipulation avec rigueur et précision  L’élève réalise un mesure de pH de la solution diluée | I1, I2, I3, I4, | |  |  |  | |
| II- a) | |  |  |  | |
| II-b) | |  |  |  | |
| **Communiquer** | L’élève reporte la mesure de pH dans le tableau  L’élève répond à la problématique de départ en justifiant sa réponse.  Jean : 🔾 A RAISON N’A PAS RAISON  mais la dilution du vinaigre :  ❑ permet de diminuer l’acidité  ❑ permet d’augmenter l’acidité  ❑ ne change pas l’acidité | I5, I6, I7 | |  |  |  | |
| II-c) | |  |  |  | |
| II-d) | |  |  |  | |
|  |  |  | |  |  |  | |

**A : acquis ECA : en cours d’acquisition NA : non acquis**

*L’activité est proposée à des élèves de première année CAP en début de formation pour une première utilisation de la grille par exemple.*

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| **La situation-problème n°1** :  ***Quelles précautions prendre avec ce produit d’usage courant ?***  On utilise dans la vie courante ce produit chimique qui permet de désinfecter, de détacher, de blanchir ou de désodoriser.  imgjavelaplati  Bouteille d'Eau de Javel (2l.)  Xi-Irritant Conseils d'utilisation : **• Pour entretien de la maison :** **Pour les sols, murs, carrelages et sanitaires** : diluer 1 verre dans 5 litres d’eau, laisser agir 5 mn puis rincer. **Pour les toilettes** : verser directement sur les parois de la cuvette 1 verre et demi d’eau de javel et laisser agir 15 mn environ, rincer. **Pour les poubelles**, **vide-ordures** : diluer 2 verres d’eau de javel pour 1 litre d’eau, laissé agir 15 mn, rincer.  **A-** Jean a effacé par erreur le nom du produit,. Indiquer le nom de ce produit :. \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  **B-** Ce produit est-il **dangereux** ? **Expliquer** pourquoi .  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  **C-** Afin de prendre en compte la signification de ce pictogramme, **cocher** les pictogrammes de précautions que l’on doit mettre pour **manipuler** ce produit.  auto0  ❑ ❑ ❑ ❑ ❑ ❑  2  http://assessfirst.blogs.com/.a/6a00d8341d444553ef0134858ff426970c-800wi***Jean prétend que ce produit est « très acide ».***  ***Problématique :***  **À votre avis, cette affirmation est-elle vraie ? Comment pourrait-on la vérifier expérimentalement ?** |
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| **HYPOTHÈSE** (Quelle est mon idée ?)  **Je pense que** \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ |  |

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| **MÉTHODE DE RÉSOLUTION** (Comment prouver ou non mon hypothèse ?)  **Je vais utiliser :** (écrire le matériel dont vous avez besoin pour essayer de prouver votre hypothèse)  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  **Je décris ma démarche :** (Indiquer toutes les actions à mettre en œuvre pour résoudre cette affirmation en commençant vos phrases par des tirets suivis de verbes à l’infinitif)  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_  **Mon hypothèse sera vérifiée si \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_\_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_**  \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ \_ |  |

**Appel n° 1 : Appeler l’examinateur, et à l’ORAL :**

* **Expliquer le problème à résoudre**
* **Proposer une méthode expérimentale qui vous permettrait de vérifier ou non l’affirmation de l’élève.**

1. **Mesure du pH d’une solution commerciale d’eau de javel avec pH-mètre stylo**

**1-** En respectant les règles de sécurité, **remplir** un bécher à moitié d’eau de Javel.

**2-** **Placer** le pH mètre stylo dans le bécher.

**3**- **Allumer** le pH-mètre stylo et le laisser se stabiliser (environ 30 à 45 s).

**4-** **Relever** la valeur du pH indiquée par le pH-mètre.

**pH Eau de Javel = …………**

**5-** **Éteindre** le pH-mètre, retirer la sonde, la rincer à l'eau du robinet puis à l'eau distillée et la sécher délicatement avec le papier absorbant.



**6-** Au vu du pH, l’eau de Javel a un caractère (**cocher** la bonne réponse):

❑ Acide ❑ Neutre ❑ Basique

**7-** L’affirmation de l’élève est-elle vraie ?



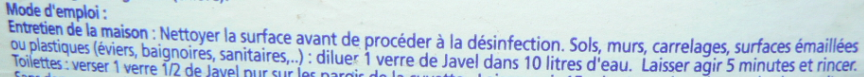
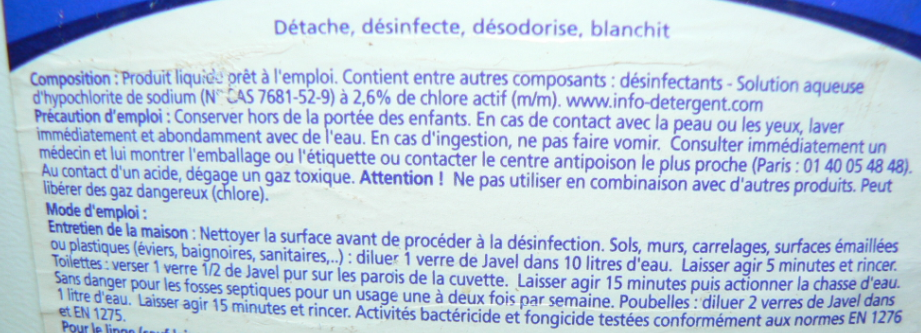
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**La situation-problème n°2 :**

Vous trouverez ci- dessous un extrait du mode d’emploi de ce produit



***Problématique :***

À partir des conseils d’utilisation de ce produit, ***que faut-il faire pour l’utiliser comme produit d’entretien à la maison ? Pour quelles raisons faire cela ?***

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**Appel n° 2 : Appeler l’examinateur:**

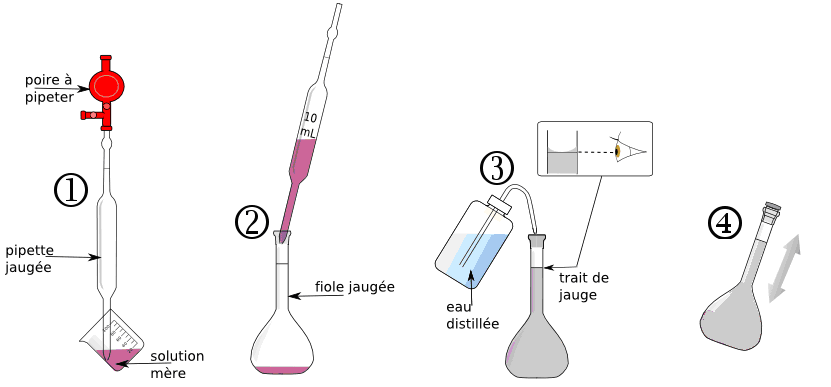
* **Réaliser dans le respect des règles de sécurité la dilution de la solution commerciale.**

1. **Évolution du pH au cours de la réalisation d’une dilution au 10ème de la solution commerciale d’eau de Javel**
2. Réalisation de la dilution

① Prélever 10 mL de la solution mère contenue dans le bécher étiqueté « solution d’eau de Javel » à l’aide d’une pipette jaugée de 10 mL, munie d’une poire à pipeter.

② Verser les 10 mL de solution initiale dans une fiole jaugée de 100 mL

③ Compléter la fiole jaugée avec de l’eau déminéralisée jusqu’au trait de jauge.

④ Fermer la fiole jaugée avec un bouchon et homogénéiser la solution.

1. Mesure du pH de la solution diluée

* Verser le contenu de la fiole jaugée dans le bécher étiqueté « préparation diluée ».
* Déterminer la valeur du pH de cette nouvelle solution à l’aide du stylo-pH.
* Noter la valeur du pH de la solution diluée : pH = …………..

1. Exploitation

Cocher les bonnes réponses à l’aide de résultats expérimentaux précédents.

La dilution de la solution d’eau de Javel :

□ ne change pas la valeur du pH

□ entraine une diminution de la valeur du pH

□ entraine une augmentation de la valeur du pH

La solution diluée :

□ est plus basique que la solution mère

□ est moins basique que la solution mère



***d) A l’aide de l’étude précédente en déduire pour quelles raisons réaliser une dilution pour ce type de produit ?***



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